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| Question 1Rusting is an example of a:1. Chemical Change
2. Chemical Property
3. Physical Change
4. Physical Property
 | Question 2Name 4 signs of a Chemical Reaction that we have discussed or seen |
| Question 3What TYPE of reaction is this?**N2 + 3H2 🡪 2NH3** | Question 4What TYPE of reaction is this?**2NaCl 🡪 2Na + Cl2** |
| Question 5What TYPE of reaction is this?**Na2S + 2HCl 🡪 2NaCl + H2S** | Question 6What TYPE of reaction is this?**2HCl + Zn 🡪 ZnCl2 + H2** |
| Question 7**The density of gold is 19g/mL**This is an example of a:1. Physical Change
2. Physical Property
3. Chemical Change
4. Chemical Property
 | Question 8**The water evaporated to steam**This is an example of a:1. Physical Change
2. Physical Property
3. Chemical Change
4. Chemical Property
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| Answer 2* Change in temperature
* Change in smell/odor
* Change in pH
* Change in color
* Gas produced (Bubbles/fizzing)
* Precipitate formed
* Light or electricity produced
 | Answer 1Chemical Change(it is an action that has made a new substance) |
| Answer 4Decomposition | Answer 3Synthesis |
| Answer 6Single Replacement(Single Displacement) | Answer 5Double Replacement(Double Displacement) |
| Answer 8Physical Change | Answer 7Physical Property |
| Question 9**The ethanol is flammable**This is an example of a:1. Physical Change
2. Physical Property
3. Chemical Change
4. Chemical Property
 | Question 10**The freezing point of water is 0oC**This is an example of a:1. Physical Change
2. Physical Property
3. Chemical Change
4. Chemical Property
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| Question 11An unknown substance tastes sour and has a low pH. Is this substance an acid, base, or neutral? | Question 12What would you need to add in order to neutralize a weak base?1. A strong base
2. A weak base
3. An acid
4. A neutral substance
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| Question 13An unknown substance feels slippery and has a high pH. Is this substance an acid, base, or neutral? | Question 14What pH level is considered neutral? |
| Question 15What is the formula for calculating density? | Question 161. What is the density of an object that has a mass of 15 g and a volume of 5 cm3?
2. Would this substance sink or float in water?
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| Answer 10Physical Property | Answer 9Chemical Property |
| Answer 12c. An acid | Answer 11An acid |
| Answer 14pH of 7 is neutral  | Answer 13A base |
| Answer 161. The density would be 3 g/cm3
2. This object would sink since its density is greater than 1
 | Answer 15Density = mass ÷ volume (or mass/volume) |
| Question 17A reaction in which a system absorbs energy from its surroundings and would feel colder is called a(n) \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ reaction. | Question 18A solid that forms and settles out of a liquid mixture during a chemical reaction is called a(n) \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| Question 19A change in a substance that does NOT cause it to change into a new substance or change many of its properties | Question 20A property of a substance that COULD permanently change the identity of the substance |
| Question 21A description of a substance that can be observed without changing the identity of the substance  | Question 22A measurement of how close or compact the particles are in a substance |
| Question 23What are 2 ways to increase the rate of a reaction? | Question 24You have 2 blocks of aluminum, one is large and one is small. How would their densities compare?1. The small one has the highest density
2. The large one has the highest density
3. Both would have the same density
4. There is no way to know without measuring them
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| Answer 18Precipitate | Answer 17Endothermic |
| Answer 20Chemical Property | Answer 19Physical change |
| Answer 22Density | Answer 21Physical Property |
| Answer 24Their densities would be the same since they are both made of the same substance (aluminum) | Answer 231. Increase temperature
2. Increase surface area
3. Stir or increase motion of particles
4. Increase concentration
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| Question 25Use the following terms to complete the sentences: **hydroxide, hydronium**Acids produce excess \_\_\_\_\_\_\_\_\_\_\_ionsBases produce excess \_\_\_\_\_\_\_\_\_\_\_ions | Question 26What range of numbers on a pH scale would indicate a substance is an ACID? |
| Question 27What TYPE of reaction is this?**C2H4  + 3O2  → 2CO2  + 2H2O** | Question 28What TYPE of reaction is this?**FeSO4  + Cu 🡪 CuSO4 + Fe** |
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| Answer 26Acids = pH less than 7 | Answer 25Acids produce excess h**ydronium** ionsBases produce excess **hydroxide** ions |
| Answer 28Single Replacement | Answer 27Combustion |